

## **Cambridge International Examinations**

Cambridge International General Certificate of Secondary Education

#### **CO-ORDINATED SCIENCES**

0654/12

Paper 1 Multiple Choice (Core)

May/June 2018

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

#### **READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

### Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.





1 Which rows correctly match characteristics of living things with their descriptions?

	characteristic description						
1	excretion	removing the waste products of metabolism					
2	growth	growth making more living things of the same type					
3	nutrition	taking in or producing food					
4	respiration	releasing energy from food					

- **A** 1, 2 and 4
- **B** 1, 3 and 4
- C 1 and 3 only
- **D** 2 and 4 only

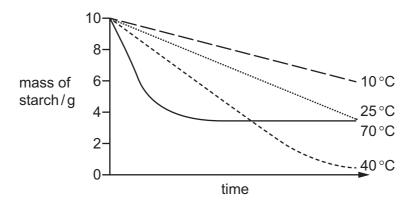
2 Which statement about cells is correct?

- A Cell membranes are found only in animal cells.
- **B** Cell membranes are found only in plant cells.
- C Cell walls are found only in animal cells.
- **D** Cell walls are found only in plant cells.

3 Which line shows the structures in increasing size?

- **A** chromosome  $\rightarrow$  gamete  $\rightarrow$  gene  $\rightarrow$  nucleus
- **B** chromosome  $\rightarrow$  nucleus  $\rightarrow$  gene  $\rightarrow$  gamete
- $\mathbf{C}$  gene  $\rightarrow$  chromosome  $\rightarrow$  gamete  $\rightarrow$  nucleus
- **D** gene  $\rightarrow$  chromosome  $\rightarrow$  nucleus  $\rightarrow$  gamete

**4** The graph shows the rate at which 10 g of starch is broken down by amylase at four temperatures.



Which is the optimum temperature?

- **A** 10 °C
- **B** 25 °C
- **C** 40 °C

0654/12/M/J/18

**D** 70 °C

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5 Tests were carried out on a colourless liquid, with the following results.

test	colour obtained		
Benedict's	blue		
biuret	purple		
iodine	blue/black		

What did the colourless liquid contain?

- A protein only
- **B** protein and reducing sugar only
- C protein and starch only
- **D** protein, reducing sugar and starch
- **6** Which statement is correct?
  - **A** The pulmonary artery carries deoxygenated blood away from the left ventricle.
  - **B** The pulmonary artery carries deoxygenated blood away from the right ventricle.
  - **C** The pulmonary vein carries oxygenated blood away from the left ventricle.
  - **D** The pulmonary vein carries oxygenated blood away from the right ventricle.
- 7 By which process does oxygen pass from the alveoli to the blood capillaries in the lungs?
  - **A** diffusion
  - **B** evaporation
  - C secretion
  - **D** transpiration
- 8 What happens when the human body temperature drops below normal?

	arterioles near skin surface	sweat secreted
Α	constrict	no
В	constrict	yes
С	dilate	no
D	dilate	yes

**9** Which row is true of asexual reproduction?

	number of parents	offspring
Α	1	genetically dissimilar
В	1	genetically identical
С	2	genetically dissimilar
D	2	genetically identical

- 10 What is a function of the stigma of a flower?
  - A to make female gametes
  - B to make male gametes
  - C to produce nectar to attract insects
  - **D** to secrete a sugary solution to aid the germination of pollen grains
- 11 In a plant, blue flower colour is dominant to red flower colour. A heterozygous blue-flowered plant is crossed with another heterozygous blue-flowered plant.

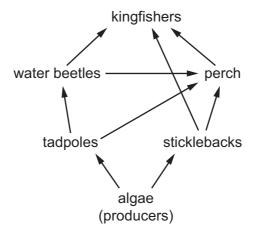
What are the expected proportions of the flower colour of the offspring?

- A 25% blue, 75% red
- **B** 50% blue, 50% red
- C 75% blue, 25% red
- **D** 100% blue, 0% red
- 12 Cows have been bred to produce much greater yields of milk than cows from a century ago.

What is this an example of?

- A artificial selection
- **B** conservation
- **C** inheritance
- **D** natural selection

13 The diagram shows a food web.



Which of the animals are carnivores?

- A kingfishers only
- **B** kingfishers, perch and water beetles
- C perch and water beetles only
- D tadpoles and sticklebacks
- **14** Pure copper chloride can be obtained from a mixture of powdered copper and solid copper chloride.

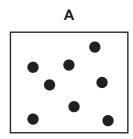
Three stages in the method are listed.

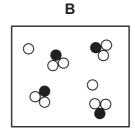
- P add water and stir
- Q crystallise
- R filter

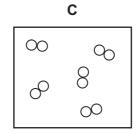
In which order are these stages carried out in order to obtain pure copper chloride from the mixture?

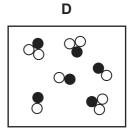
- **A**  $P \rightarrow Q \rightarrow R$
- $\mathbf{B} \quad \mathsf{P} \, \to \, \mathsf{R} \, \to \, \mathsf{Q}$
- $\mathbf{C} \quad \mathsf{R} \to \mathsf{P} \to \mathsf{Q}$
- $\mathbf{D} \quad \mathsf{R} \, \to \, \mathsf{Q} \, \to \, \mathsf{P}$

15 Which diagram represents a mixture of an element and a compound?

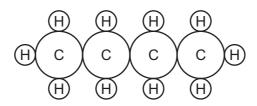








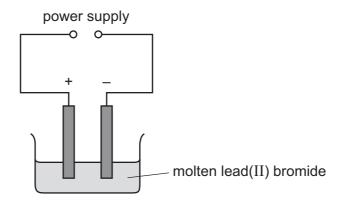
**16** The diagram represents a molecule of butane.



What is the formula of butane?

- **A** C<sub>2</sub>H<sub>5</sub>
- $\mathbf{B}$   $C_4H_8$
- $\boldsymbol{C} \quad C_4 H_{10}$
- **D**  $C_{10}H_4$

**17** Molten lead(II) bromide is electrolysed as shown.



An element is produced at the negative electrode.

What is the name of the element and of the negative electrode?

	element	negative electrode
Α	bromine	anode
В	bromine	cathode
С	lead	anode
D	lead	cathode

- 18 Which statement about electroplating iron with chromium is correct?
  - A A catalyst is used.
  - **B** The anode is chromium.
  - **C** The electrolyte contains aqueous iron ions.
  - **D** The electrolyte contains solid chromium ions.
- 19 Calcium carbonate reacts with dilute hydrochloric acid.

Equal masses of different-sized pieces of calcium carbonate are placed in four test-tubes, as shown.

test-tube	1	2	3	4
size of calcium carbonate	medium pieces	powder	small pieces	large pieces

Equal volumes of the same concentration of dilute hydrochloric acid are added to each test-tube.

Which test-tube shows the lowest rate of reaction?

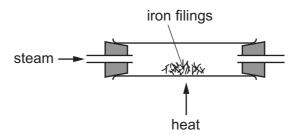
**A** 1

**B** 2

C (

D 4

20 When iron is heated with steam, a black solid is formed.



The equation for the reaction is shown.

Which statement about this reaction is correct?

- A Iron has been oxidised because it has gained oxygen.
- **B** Iron has been reduced because it removed oxygen from water.
- **C** Iron oxide has been reduced because it contains oxygen.
- **D** Water has been oxidised because it contains oxygen.

21 Magnesium and hydrochloric acid react together.

What is the correct word equation?

- A magnesium + hydrochloric acid → magnesium chloride + hydrogen
- **B** magnesium + hydrochloric acid → magnesium chloride + hydrogen chloride
- **C** magnesium + hydrochloric acid → magnesium chloride + hydrogen + chlorine
- **D** magnesium + hydrochloric acid → magnesium chloride + hydrogen chloride + hydrogen
- 22 Some properties of elements are listed.
  - 1 conduct electricity
  - 2 form coloured compounds
  - 3 high boiling point

What are the properties of a transition element?

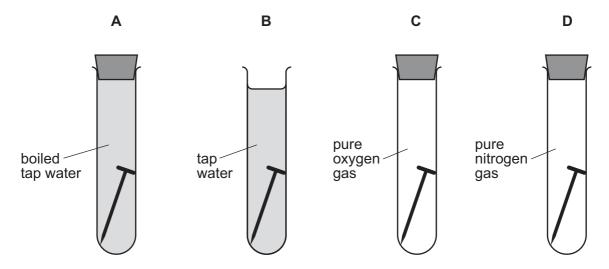
- **A** 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3
- **23** Metal X reacts rapidly with steam but only very slowly with cold water.

What is X?

- A calcium
- **B** copper
- **C** magnesium
- **D** sodium
- **24** Which gas is **not** a common air pollutant?
  - A water vapour
  - **B** carbon monoxide
  - C nitrogen dioxide
  - **D** sulfur dioxide

25 Four iron nails are placed in four test-tubes as shown.

In which test-tube does the iron nail rust most quickly?



26 Calcium carbonate is decomposed by heating in an industrial process.

The equation for this reaction is shown.

calcium carbonate → calcium oxide + carbon dioxide

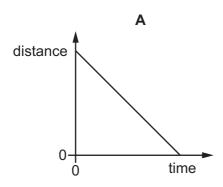
Which statement is not correct?

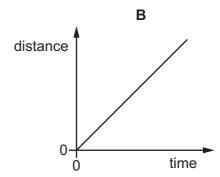
- **A** The common name for calcium carbonate is limestone.
- **B** The common name for calcium oxide is lime.
- **C** Calcium oxide is used to neutralise alkaline soil.
- **D** Calcium oxide is used to neutralise industrial waste products.
- **27** Ethene is used to make poly(ethene).

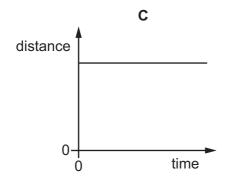
Which words describe ethene?

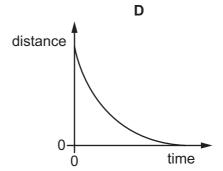
- 1 hydrocarbon
- 2 saturated
- 3 monomer
- **A** 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

28 Which distance-time graph represents a body moving with a changing speed?



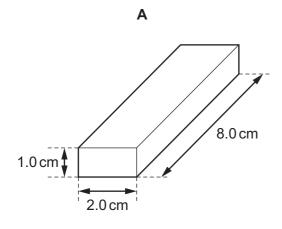


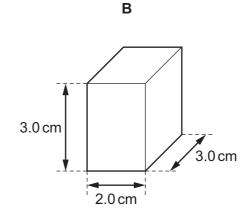


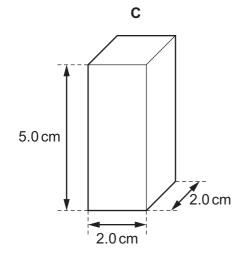


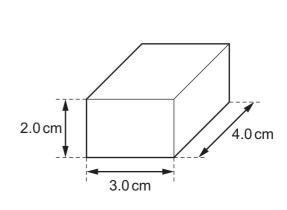
**29** The diagrams show four solid blocks with the same mass.

Which block is made from the least dense material?









D

**30** A boy carries out an experiment to demonstrate pressure and its relationship to force and area.

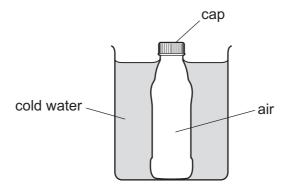
Which experiment produces the highest pressure on the classroom floor?

- A standing with one foot on the floor
- **B** standing with two feet on the floor
- **C** standing with one foot on the floor, holding a 5.0 kg mass
- **D** standing with two feet on the floor, holding a 5.0 kg mass

31 Which energy resource does **not** use a turbine and generator to produce electricity?

- A geothermal
- **B** nuclear fission
- C solar cells
- **D** wind

**32** A glass bottle containing warm air is sealed with a screw cap and then cooled in cold water.



The contraction of the glass bottle can be ignored.

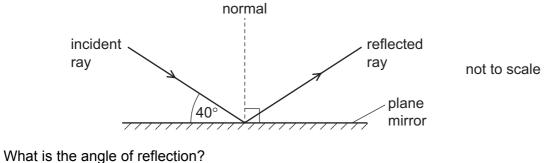
What remains the same during the cooling?

- A the air pressure inside the bottle
- **B** the energy of the air molecules in the bottle
- C the force on the cap made by the air molecules in the bottle
- **D** the volume of air in the bottle
- 33 Which type of heat transfer is the main method in liquid water?
  - A conduction
  - **B** convection
  - **C** evaporation
  - **D** radiation
- **34** A girl is sitting on a rock in the sea looking at passing waves. She notices that five complete wavelengths pass her in 20 s.

What is the frequency of this wave?

- **A** 0.25 Hz
- **B** 4.0 Hz
- **C** 15 Hz
- **D** 100 Hz

**35** The diagram shows light hitting a plane mirror.



40°

**B** 50°

80°

100°

**36** White light is dispersed by a glass prism into the colours of the spectrum.

Which colour of light is refracted the most and which is refracted the least?

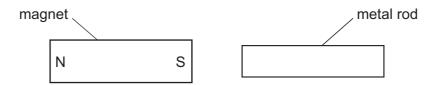
	refracted most	refracted least
Α	green	red
В	red	green
С	red	violet
D	violet	red

**37** The electromagnetic spectrum includes radio waves, infra-red and X-rays.

What is the correct sequence of these waves in order of increasing wavelength (smallest wavelength first)?

- infra-red, radio waves, X-rays
- infra-red, X-rays, radio waves
- X-rays, infra-red, radio waves C
- X-rays, radio waves, infra-red

**38** A bar magnet is brought near to a metal rod.



The magnet is now turned around so that the N-pole is on the right. The magnet is again brought near to the metal rod.

In both cases the metal rod is attracted to the magnet.

What could the metal rod be?

- **A** another bar magnet
- B a piece of aluminium
- C a piece of copper
- **D** a piece of iron

**39** Which row correctly states whether the unit for electromotive force (e.m.f.), mass and weight is the newton?

	electromotive force (e.m.f.)	mass	weight
Α	no	no	yes
В	no	yes	yes
С	yes	no	no
D	yes	yes	no

- **40** Which changes **both** result in an increase in the resistance of a metal wire?
  - A decreasing the length and decreasing the diameter of the wire
  - **B** decreasing the length and increasing the diameter of the wire
  - **C** increasing the length and decreasing the diameter of the wire
  - **D** increasing the length and increasing the diameter of the wire

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The Periodic Table of Elements

	<b> </b>	ه ح لا	helium 4	10	Ne	neon 20	18	Ā	argon 40	36	첫	krypton 84	54	Xe	xenon 131	98	牊	radon			
	=			6	ш	fluorine 19	17	Cl	chlorine 35.5	35	ğ	bromine 80	53	П	iodine 127	85	¥	astatine -			
	>			8	0	oxygen 16	16	ഗ	sulfur 32	34	Se	selenium 79	52	<u>е</u>	tellurium 128	84	Ъо	polonium -	116		livemorium –
	>			7	z	nitrogen 14	15	۵	phosphorus 31	33	As	arsenic 75	51	Sp	antimony 122	83	Ξ	bismuth 209			
	≥			9	ပ	carbon 12	14	S	silicon 28	32	Ge	germanium 73	20	Sn	tin 119	82	Pb	lead 207	114	Εl	flerovium
	≡			2	М	boron 11	13	Αl	aluminium 27	31	Ga	gallium 70	49	In	indium 115	81	lΤ	thallium 204			
							•			30	Zn	zinc 65	48	В	cadmium 112	80	Рд	mercury 201	112	ر ت	copernicium
										29	Cn	copper 64	47	Ag	silver 108	62	Αn	gold 197	111	Rg	roentgenium -
Group										28	Z	nickel 59	46	Pq	palladium 106	78	చ	platinum 195	110	Ds	darmstadtium -
วือ				-						27	ဝိ	cobalt 59	45	뫈	rhodium 103	77	'n	iridium 192	109	¥	meitnerium -
		- I	hydrogen 1							56	Fe	iron 56	44	R	ruthenium 101	92	SO	osmium 190	108	Hs	hassium -
							,			25	Mn	manganese 55	43	ည	technetium -	75	Re	rhenium 186	107	Bh	bohrium —
				_	pol	ass				24	ပ်	chromium 52	42	Mo	molybdenum 96	74	≥	tungsten 184	106	Sg	seaborgium -
			Key	atomic number	atomic symbo	name relative atomic mass				23	>	vanadium 51	41	q	niobium 93	73	<u>Б</u>	tantalum 181	105	op O	dubnium -
					atc	re				22	j=	titanium 48	40	Zr	zirconium 91	72	Ξ	hafnium 178	104	꿉	rutherfordium -
										21	လွ	scandium 45	39	>	yttrium 89	57-71	lanthanoids		89–103	actinoids	
	=			4	Be	beryllium 9	12	Mg	magnesium 24	20	Ca	calcium 40	38	Š	strontium 88	56	Ba	barium 137	88	Ra	radium
	_			8	:=	lithium 7	7	Na	sodium 23	19	$\prec$	potassium 39	37	Rb	rubidium 85	55	S	caesium 133	87	ъ́	francium -

7.1	Γn	lutetium	175	103	۲	lawrencium	I
70	Υp	ytterbium	173	102	%	nobelium	I
69	Tm	thulium	169	101	Md	mendelevium	I
89	Щ	erbium	167	100	Fm	ferminm	I
29	웃	holmium	165	66	Es	einsteinium	I
99	ρ	dysprosium	163	86	ర్	califomium	I
65	Тр	terbium	159	26	益	berkelium	-
64	Вd	gadolinium	157	96	Cm	curium	I
63	Ш	europium	152	98	Am	americium	_
62	Sm	samarium	150	94	Pu	plutonium	_
61	Pm	promethium	1	93	dΝ	neptunium	_
09	ρN	neodymium	144	92	$\supset$	uranium	238
69	Ā	praseodymium	141	91	Ра	protactinium	231
28	Ce	cerium	140	06	Ч	thorium	232
22	Га	lanthanum	139	88	Ac	actinium	I

lanthanoids

actinoids

The volume of one mole of any gas is  $24\,dm^3$  at room temperature and pressure (r.t.p.).